(a) > 1  (b) < 1  (c) > 1  (d) > 1  (e) > 1
sp\(^3\) N versus sp\(^3\) O
N is less electronegative; it more readily bonds and better stabilizes a (+) charge

sp\(^2\) O with 2 sources of delocalization stabilization versus sp\(^2\) O and sp\(^2\) N with no stabilization

sp\(^3\) N versus sp\(^2\) O versus sp\(^2\) N
N less electronegative than O and sp\(^3\) less electronegative than sp\(^2\)

sp N with delocalization stabilization versus sp\(^2\) O with no additional stabilization (and where the delocalized contributor from the O atom leads to (+) and makes the localized nbe pair far less basic

sp\(^3\) O versus sp\(^2\) O; sp\(^3\) less electronegative than sp\(^2\)